**Supplemental notes for Lecture 16**

Draw out molecules in the table to verify ED geometry and molecular shape for yourself. Anytime you answer a HW question, ask yourself about the ED geometry, molecular geometry, bond angles, polarity, hybridization, etc.—even if the question is NOT asking for this. Practice is the only way to get these ideas in your gut, so you will nail it on the exam.

![Lewis structures](image)

- in plane of paper
- coming toward you (wrt paper plane)
- going away from you (wrt paper plane)

What effect do lone pairs or multiple bonds have on the bond angles, and why?

Don’t forget that bent can be either ~120° or ~109°, depending on ED geometry!!! Watch out for this, and recognize which situation applies.